

Use of Carbon Dots Synthesized from Citrate as a Fluorescent Probe for Quercetin Determination in Tea and Beer Samples

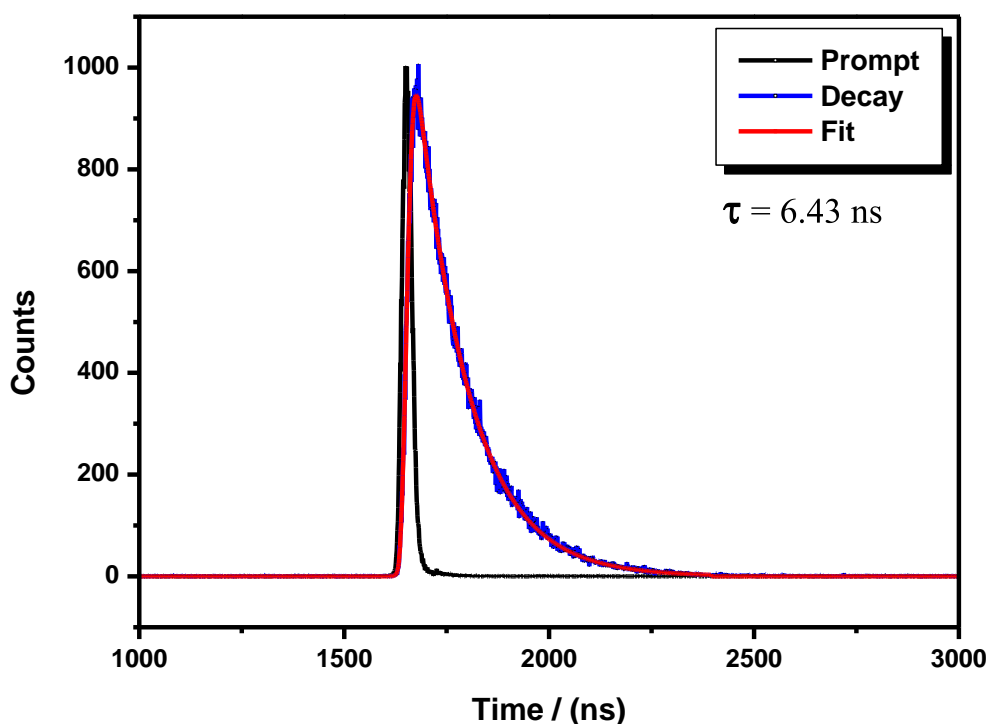
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Figure S1. Fluorescence decay profile of the CDs in the absence of quercetin in buffer solution pH 11 ($\lambda_{\text{ex}} = 380$ nm, $\lambda_{\text{em}} = 450$ nm).

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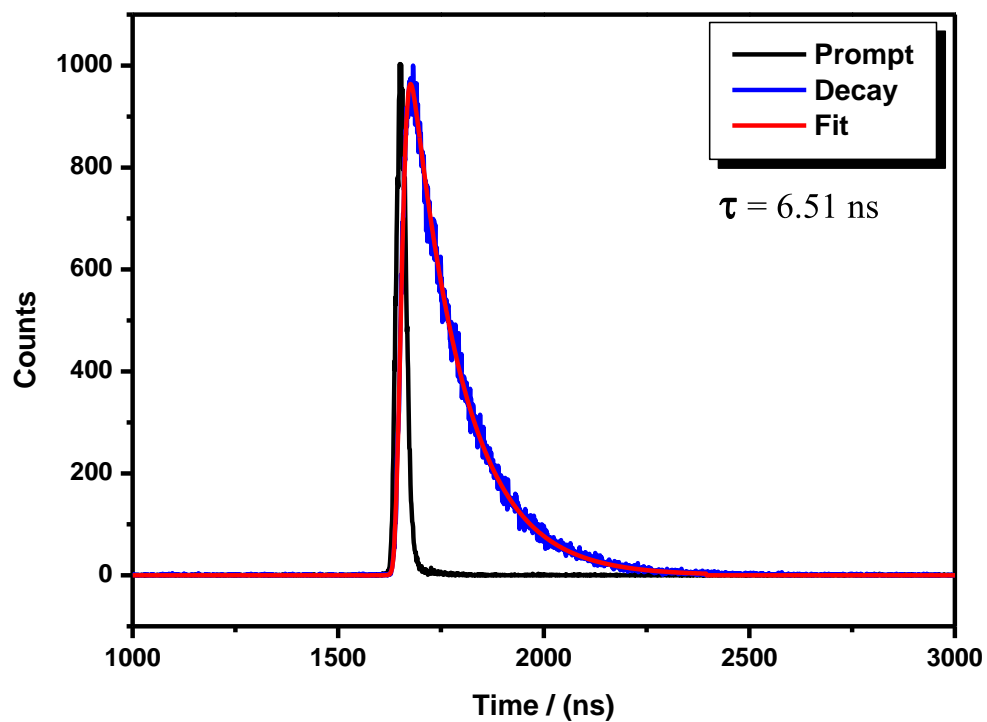


Figure S2. Fluorescence decay profile of the CDs in the presence of 1 mg L⁻¹ quercetin in buffer solution pH 11 ($\lambda_{\text{ex}} = 380$ nm, $\lambda_{\text{em}} = 450$ nm).

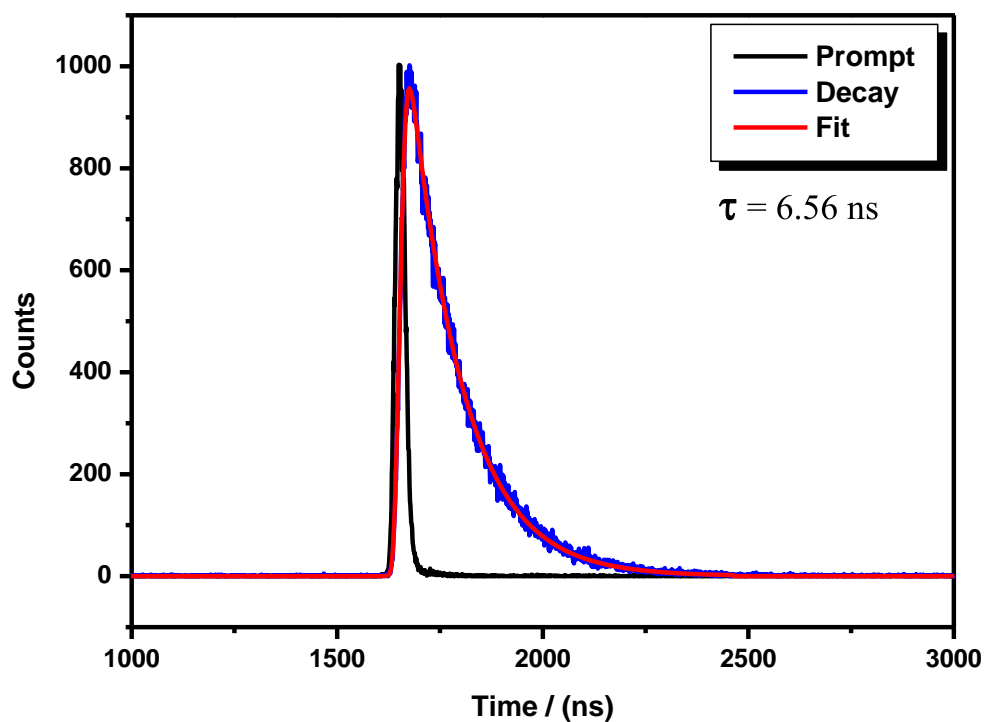


Figure S3. Fluorescence decay profile of the CDs in the presence of 4 mg L⁻¹ quercetin in buffer solution pH 11 ($\lambda_{\text{ex}} = 380$ nm, $\lambda_{\text{em}} = 450$ nm).

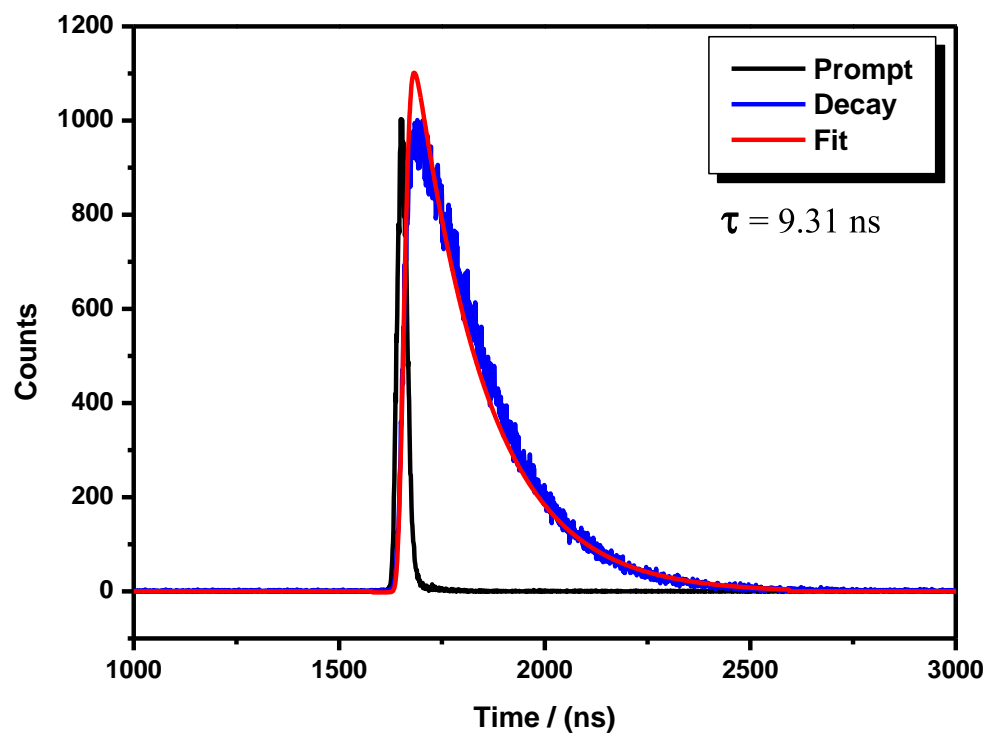


Figure S4. Fluorescence decay profile of the CDs in the aqueous solution ($\lambda_{\text{ex}} = 380 \text{ nm}$, $\lambda_{\text{em}} = 450 \text{ nm}$).

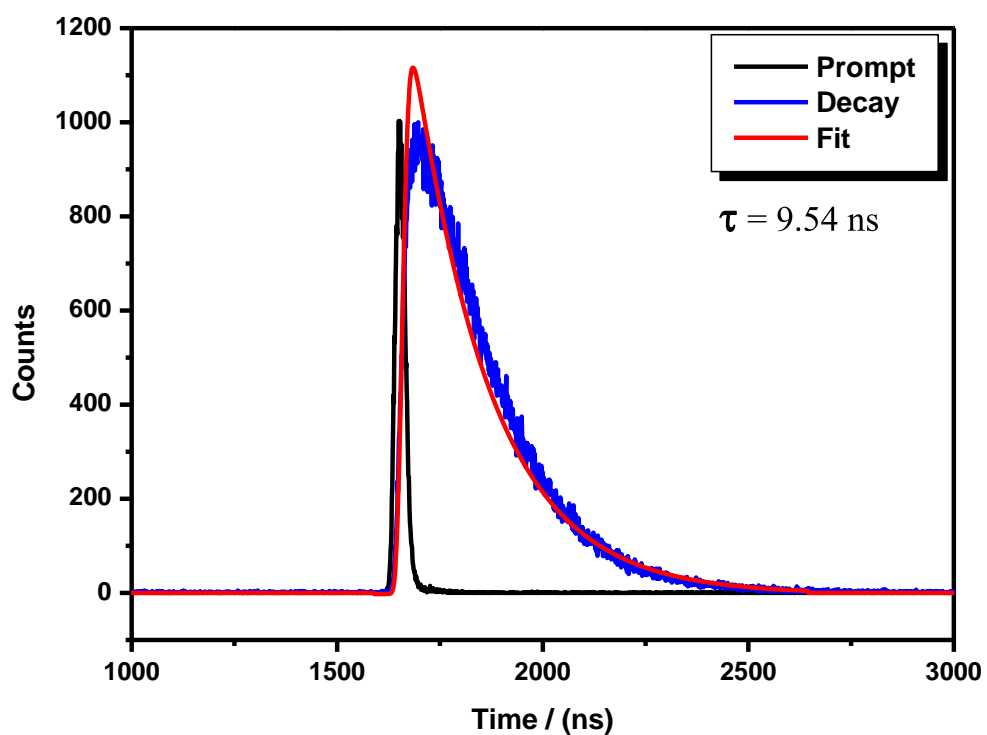


Figure S5. Fluorescence decay profile of the CDs in the presence of 4 mg L⁻¹ quercetin in the aqueous solution ($\lambda_{\text{ex}} = 380 \text{ nm}$, $\lambda_{\text{em}} = 450 \text{ nm}$).

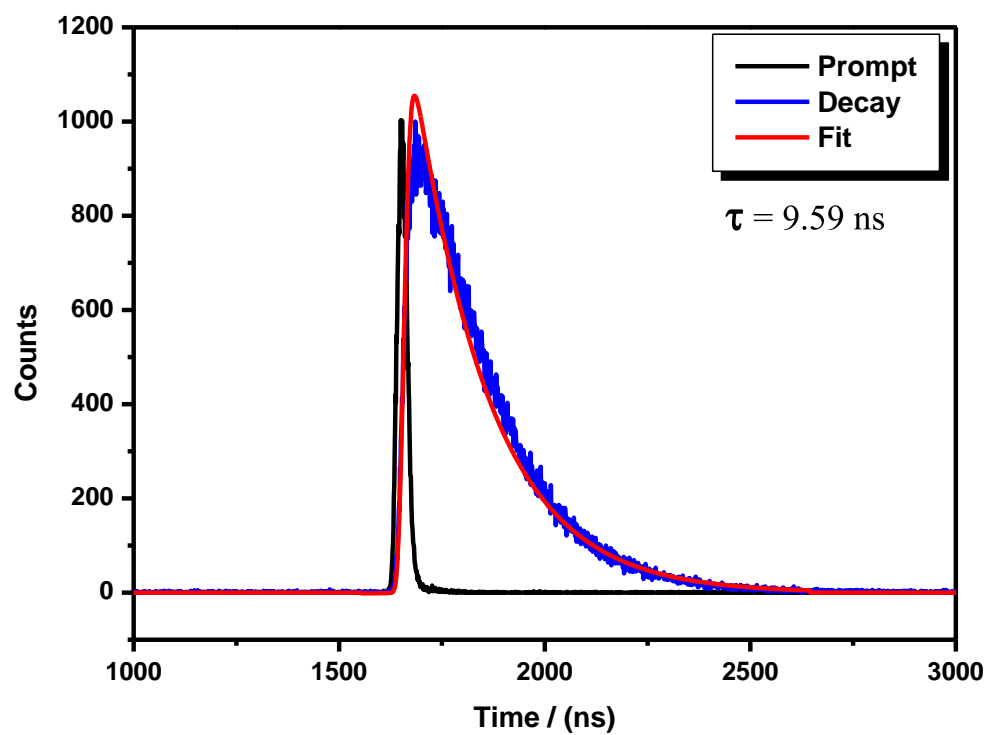


Figure S6. Fluorescence decay profile of the CDs in the presence of 10 mg L⁻¹ quercetin in the aqueous solution ($\lambda_{\text{ex}} = 380$ nm, $\lambda_{\text{em}} = 450$ nm).